Name _____ Date _____ Score _____

Free sample

Chemical Bonding

Six questions to try. Then check your answers on page 2.

Remember

Ionic: a metal gives electrons to a non-metal. **Covalent**: non-metals share pairs. **Metallic**: ions in a sea of electrons. In a formula, charges must balance.

1 What is the charge on a sodium ion?

2 Ionic or covalent: **CH_4**?

3 Write the formula: Mg^{2+} and Cl^{-}.

4 What is shared in a covalent bond?

5 Does solid salt conduct electricity? Why?

6 Why do metals conduct electricity?

Name _____ Date _____ Score _____

Answers

1. **+1 (Na^+)** — Sodium is in group 1, so it loses 1 electron.
2. **covalent** — Carbon and hydrogen are both non-metals.
3. **$MgCl_2$** — +2 is balanced by two -1 ions.
4. **a pair of electrons** — One from each atom.
5. **No; its ions cannot move.** — Molten or dissolved salt does conduct.
6. **They have delocalised electrons that can move.** — The electrons carry the charge.

In the full pack

16 questions in 5 parts (ions, ionic formulas, covalent bonding, structure and properties, and an evidence task), a notes page, a full answer key, a 14-slide teaching deck with a conductivity practical, and a 22-question online practice set (10 questions are free online).

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